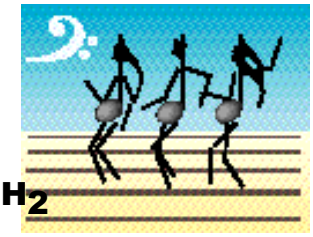
October 31, 2018
AP Chem 2-1-3
Objective: Lab Empirical and Hydrate formulas
Warm UP:



$$2Mg + 2CuSO_4 + H_2O \longrightarrow 2MgSO_4 + Cu_2O + H_2$$

1. **If 1.46g of solid magnesium are added to 500 mL of a 0.2M soln of copper (II) sulfate, what is the maximum molar yield of hydrogen gas?**

2. **When the entire limiting reagent has been consumed, how many moles of the other reagent remain?**

3. **What is the mass of copper (I) oxide in #1?**

4. **What is the conc of magnesium ions in soln at the end of the experiment?**

A solution of tin(II) nitrate is added to a solution of silver nitrate. Why is tin often used to coat steel?

Assignment:. Worksheet Practice Stoichiometry due Thursday, 2-1-4. Lab Empirical/Molecular and Hydrate on Wednesday, October 31 due Nov. 9. Lab Redox on Nov. 7-9, due Nov. 16. Lab Hard Water Nov. 14- November 16 due Nov. 20. Study Session Mondays, 7:30.