

August 29, 2018
AP Chemistry 1-1-3
Objective: Lab Chemical Reactions Day 1
Warm up:



For each of the following three reactions, first (a) write a balanced equation for the reaction and in part (b) answer the question about the reaction. Initially the coefficients should be in terms of the lowest whole numbers. Assume that solutions are aqueous unless otherwise indicated. Represent substances in solutions as ions if the substances are extensively ionized. Omit formulas for any ions or molecules that are unchanged by the reaction.

1. a. Butane is burned in air.

b. If 2 moles of each reactant is mixed, how many moles of product(s) will be produced at STP? Assume the reaction goes to completion.
2. a. A solution of sodium hydroxide is added to acetic acid.

b. What common covalent product is formed?
3. a. Solutions of barium chloride and sodium sulfate are mixed.

b. Which ion is always soluble?
4. a. Copper metal is added to concentrated nitric acid.

b. Give three observations that would occur with this reaction.

Assignment: Lab Chemical Reactions Wed. – Friday, due Friday, September 7. Notes chapter 6 due September 10. Don't forget about chapter problems!